



**Cambridge International Examinations**  
Cambridge International General Certificate of Secondary Education

**CHEMISTRY**

**0620/21**

Paper 2 Multiple Choice (Extended)

**October/November 2016**

**45 minutes**

Additional Materials:      Multiple Choice Answer Sheet  
   Soft clean eraser  
   Soft pencil (type B or HB is recommended)

\*  
6  
9  
4  
4  
2  
4  
4  
7  
5  
0  
4  
1  
\*

**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

**DO NOT WRITE IN ANY BARCODES.**

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **20** printed pages.

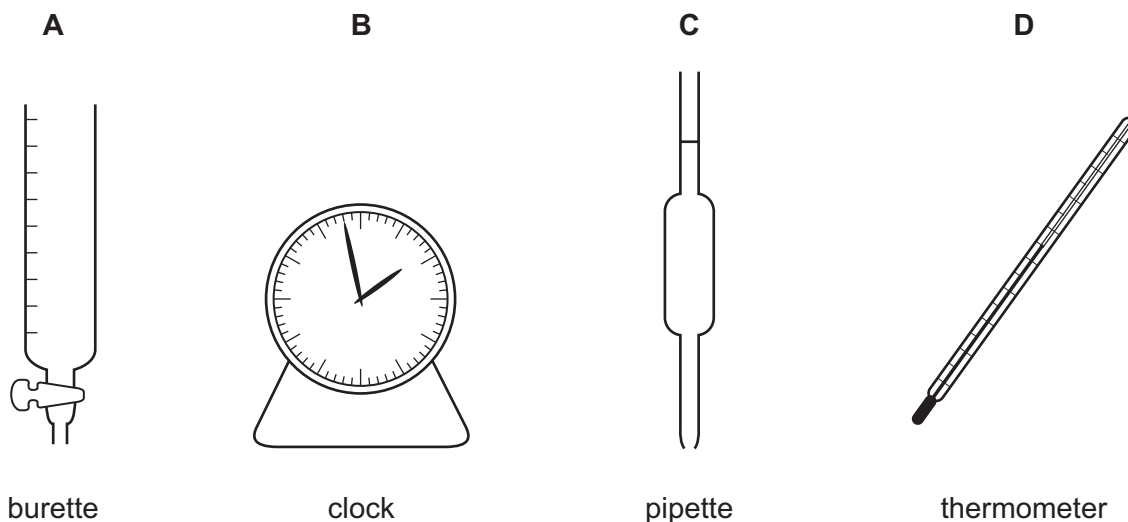
- 1 'Particles moving **very slowly** from an area of higher concentration to an area of lower concentration.'

Which process is being described?

- A a liquid being frozen
  - B a solid melting
  - C a substance diffusing through a liquid
  - D a substance diffusing through the air
- 2 A student mixes  $25\text{cm}^3$  samples of dilute hydrochloric acid with different volumes of aqueous sodium hydroxide.

In each case, the student measures the change in temperature to test if the reaction is exothermic.

Which piece of apparatus is **not** needed?



- 3 Information about the solubility of four solids, P, Q, R and S, is given in the table.

	P	Q	R	S
solubility in water	dissolves	insoluble	insoluble	dissolves

A student attempted to separate mixtures of these solids using the following method.

- 1 Add the mixture to a beaker of water and stir.
- 2 Filter the mixture.
- 3 Crystallise one of the solids from the filtrate.

Which of the following mixtures could **not** be separated by this method?

- A** a mixture of P and R
- B** a mixture of Q and P
- C** a mixture of Q and R
- D** a mixture of R and S

- 4 The table shows information about atoms of three different elements.

element	proton number	nucleon number	number of protons	number of neutrons	number of electrons
chlorine	17	35	17	W	17
chlorine	17	X	17	19	17
argon	Y	40	18	22	18
potassium	19	39	19	20	Z

What are the values of W, X, Y and Z?

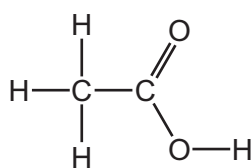
	W	X	Y	Z
<b>A</b>	18	35	18	19
<b>B</b>	18	36	18	19
<b>C</b>	19	35	19	18
<b>D</b>	19	36	19	18

5 Metal P reacts with non-metal Q to form a compound.

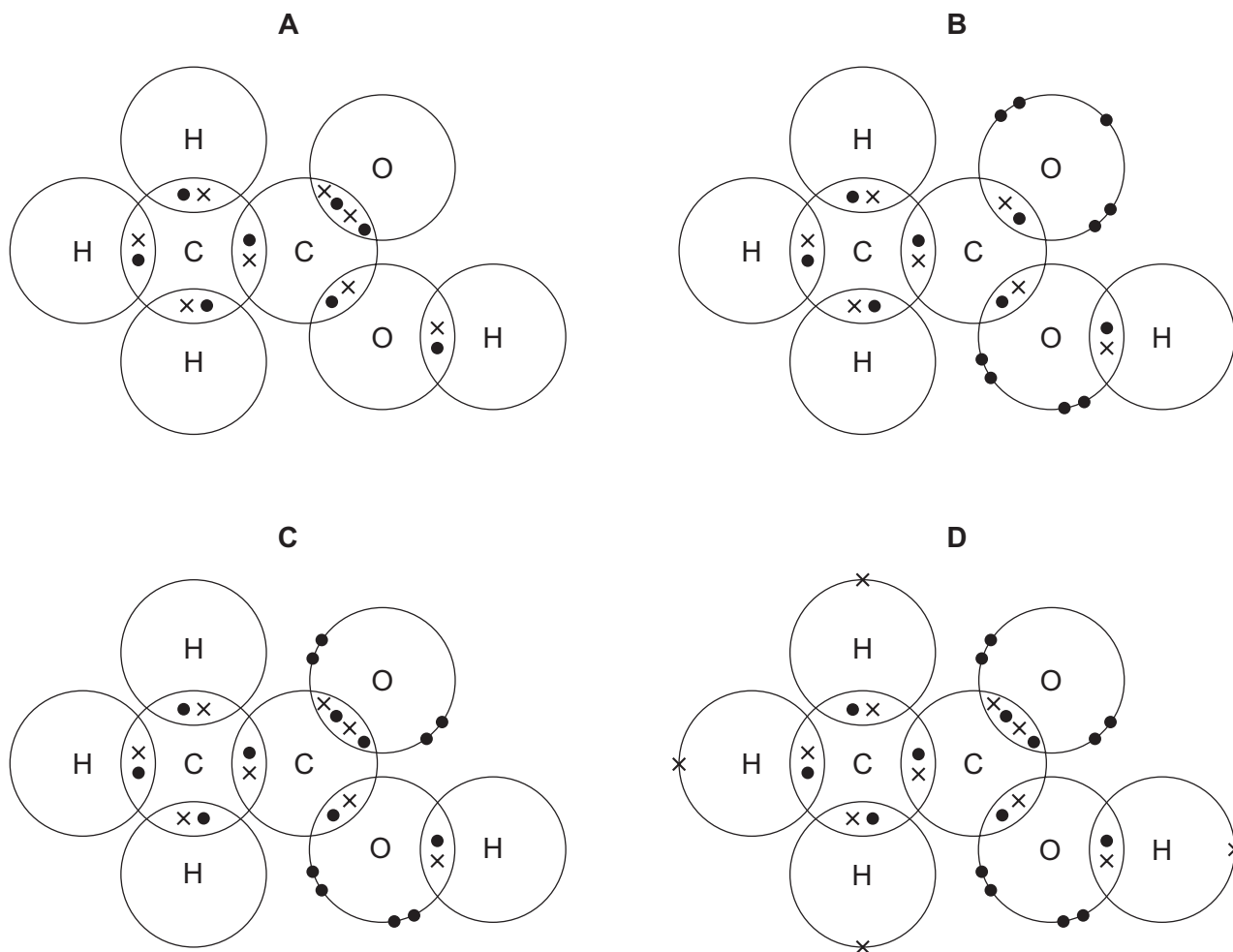
Which process takes place and which type of compound is formed?

	process	type of compound
<b>A</b>	electrons are transferred from P to Q	covalent
<b>B</b>	electrons are transferred from P to Q	ionic
<b>C</b>	electrons are transferred from Q to P	covalent
<b>D</b>	electrons are transferred from Q to P	ionic

6 The structure of ethanoic acid is shown.



Which diagram shows the arrangement of outer shell electrons in a molecule of ethanoic acid?

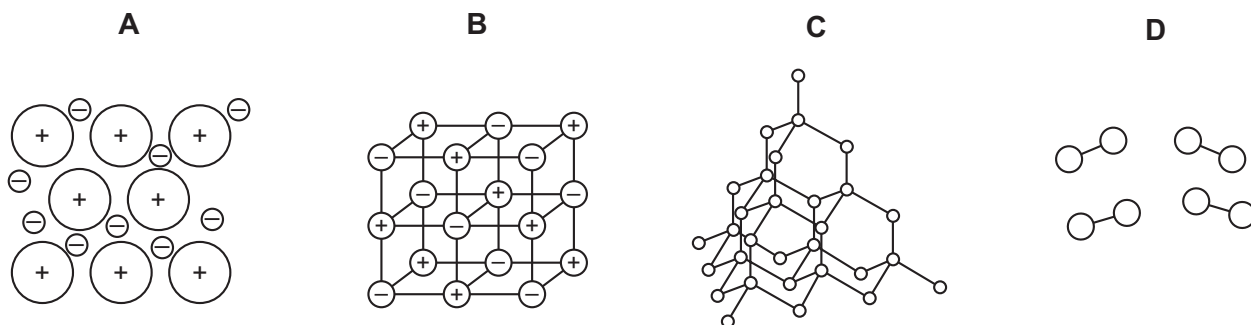


7 X is a solid at room temperature.

X has a high melting point.

Solid X conducts electricity.

Which diagram shows how the particles are arranged in solid X?



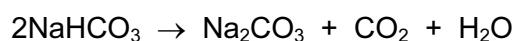
8 Benzene is a liquid with molecular formula  $C_6H_6$ .

Ethene is a gas with molecular formula  $C_2H_4$ .

Which statement is correct?

- A** 1 mole of benzene and 1 mole of ethene contain the same number of atoms.
- B** 1 mole of benzene and 1 mole of ethene both have a volume of  $24 \text{ dm}^3$  at room temperature and pressure.
- C** Both benzene and ethene have the same empirical formula.
- D** The number of carbon atoms in 0.5 moles of ethene is equal to the Avogadro constant.

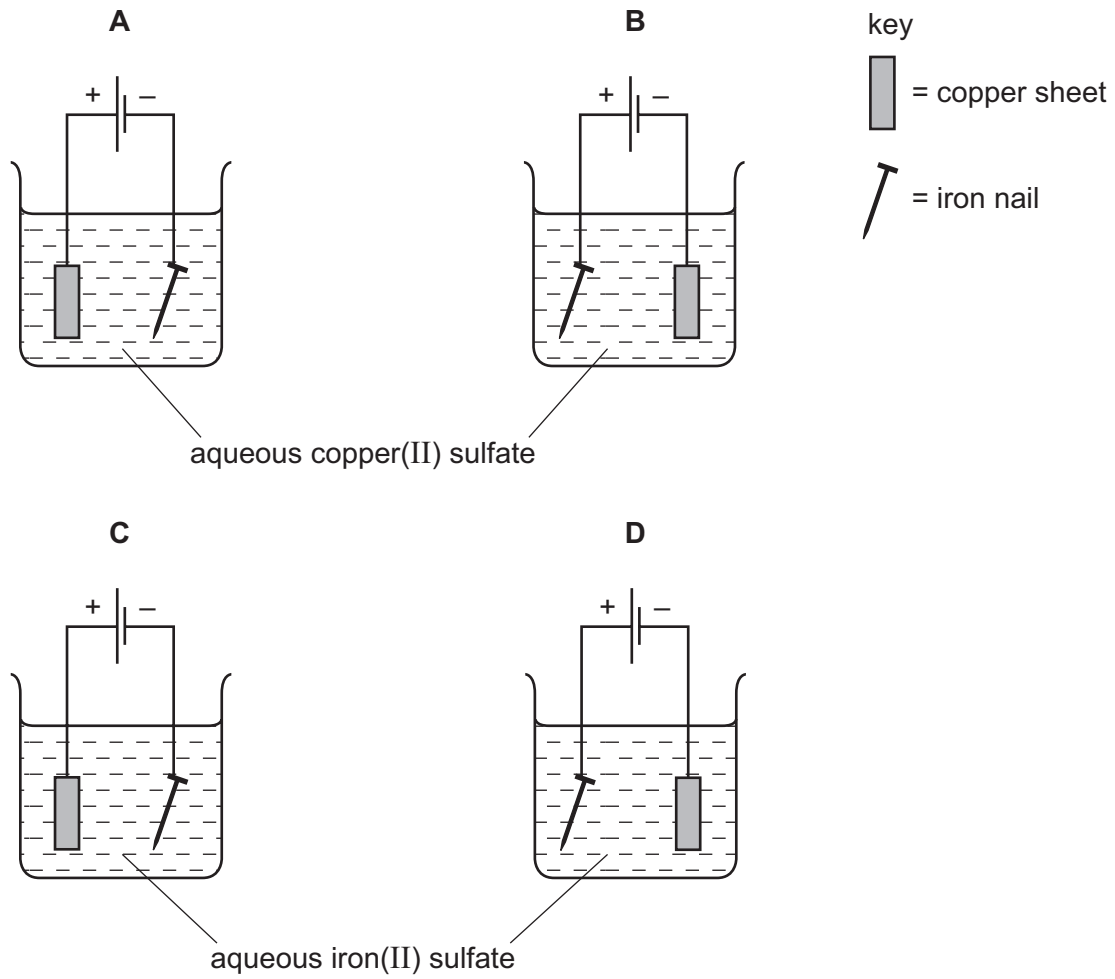
9 Sodium hydrogencarbonate undergoes thermal decomposition as shown.



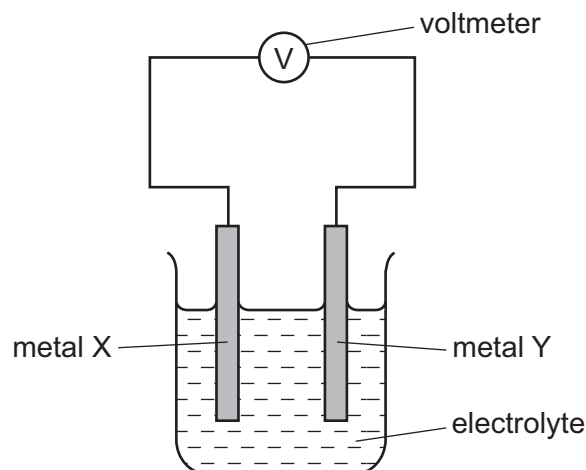
What is the maximum mass of sodium carbonate that can be made from 0.100 moles of sodium hydrogencarbonate?

- A** 4.15g
- B** 5.30g
- C** 10.6g
- D** 21.2g

10 Which apparatus could be used to electroplate an iron nail with copper?



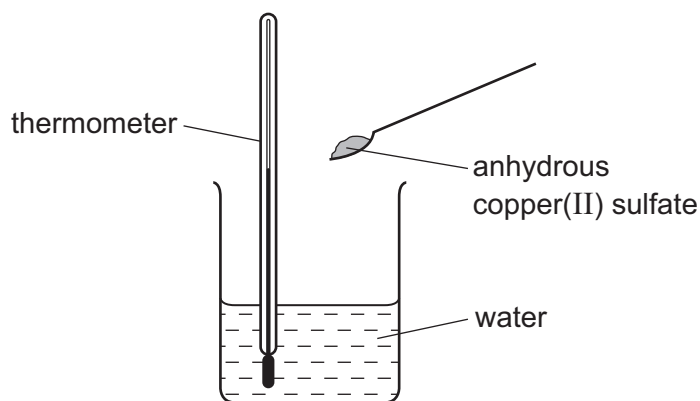
11 The diagram shows a simple cell.



Which two metals produce the highest reading on the voltmeter?

	X	Y
<b>A</b>	magnesium	copper
<b>B</b>	magnesium	iron
<b>C</b>	zinc	copper
<b>D</b>	zinc	iron

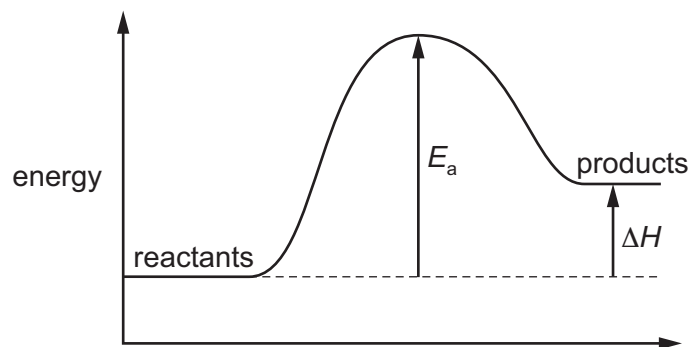
12 When anhydrous copper(II) sulfate is added to water a solution is formed and heat is given out.



Which row shows the temperature change and the type of reaction taking place?

	temperature change	type of reaction
<b>A</b>	decrease	endothermic
<b>B</b>	decrease	exothermic
<b>C</b>	increase	endothermic
<b>D</b>	increase	exothermic

13 The energy level diagram for a reaction is shown.



Which statement is **not** correct for this energy level diagram?

- A It could be the energy level diagram for the reaction when petrol is burnt.
- B Less energy is released in bond forming than is needed for bond breaking.
- C The activation energy,  $E_a$ , has a positive value.
- D The energy change,  $\Delta H$ , for the reaction is positive.

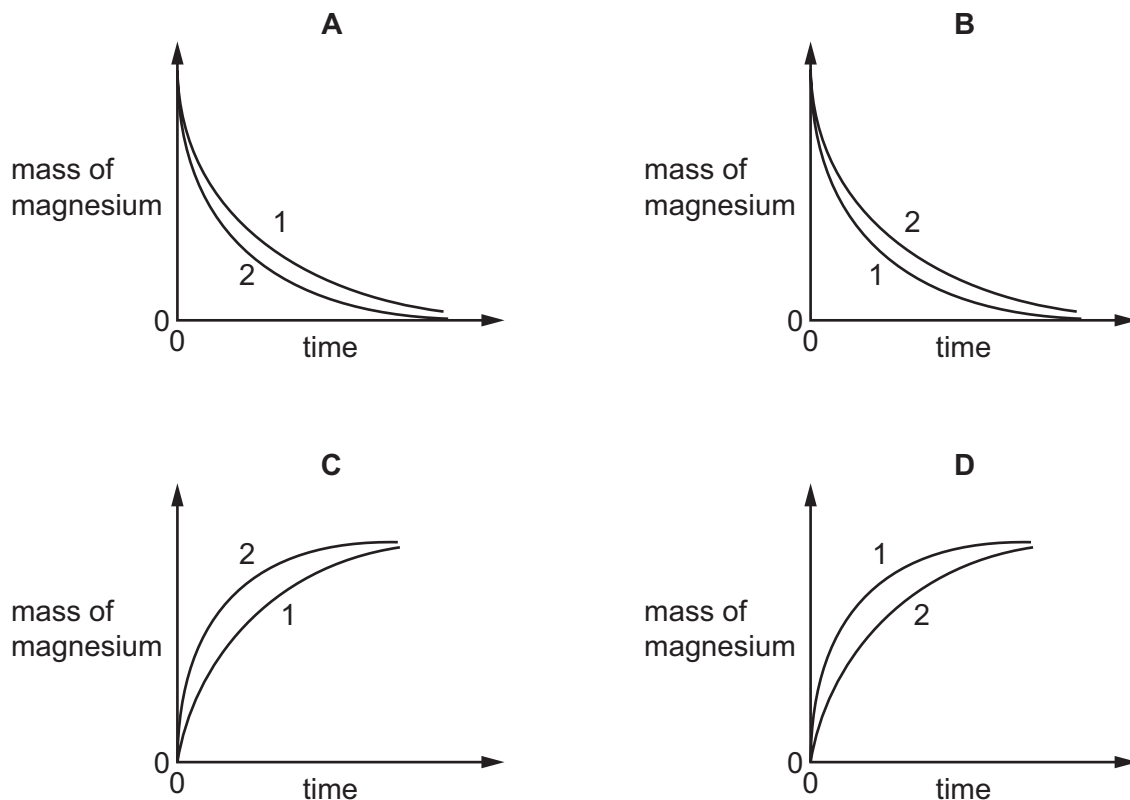


- 14** The rate of reaction between magnesium and excess dilute hydrochloric acid was followed by measuring the mass of magnesium present at regular time intervals.

Two experiments were performed.

Both experiments used 0.1g of magnesium ribbon. The acid in experiment 1 was less concentrated than in experiment 2.

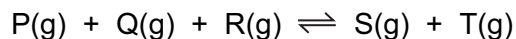
Which graph shows the results of the experiments?



- 15** Which statement explains why coal dust forms an explosive mixture with air?

- A** Coal dust catalyses the explosion.
- B** Coal dust has a large surface area.
- C** Crushing coal increases the concentration of the coal.
- D** Crushing coal increases the temperature of the coal.

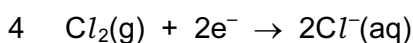
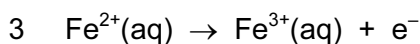
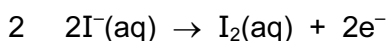
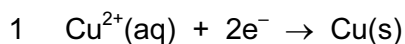
- 16 The following reversible reaction takes place in a closed vessel at constant temperature.



When the system has reached equilibrium, more T is added.

After the addition of T, which substances increase in concentration?

- A P, Q, R and S
  - B P and Q only
  - C P, Q and R only
  - D S only
- 17 Four ionic half-equations are shown.



Which statement is correct?

- A In equation 1, copper(II) ions are oxidised to copper.
  - B In equation 2, iodide ions are reduced to iodine.
  - C In equation 3, iron(II) ions are oxidised to iron(III) ions.
  - D In equation 4, chlorine is oxidised to chloride ions.
- 18 Germanium oxide is a white powder.

Germanium oxide reacts with concentrated hydrochloric acid.

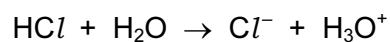
Germanium oxide reacts with concentrated aqueous sodium hydroxide.

Germanium oxide does not dissolve when added to water.

Which type of oxide is germanium oxide?

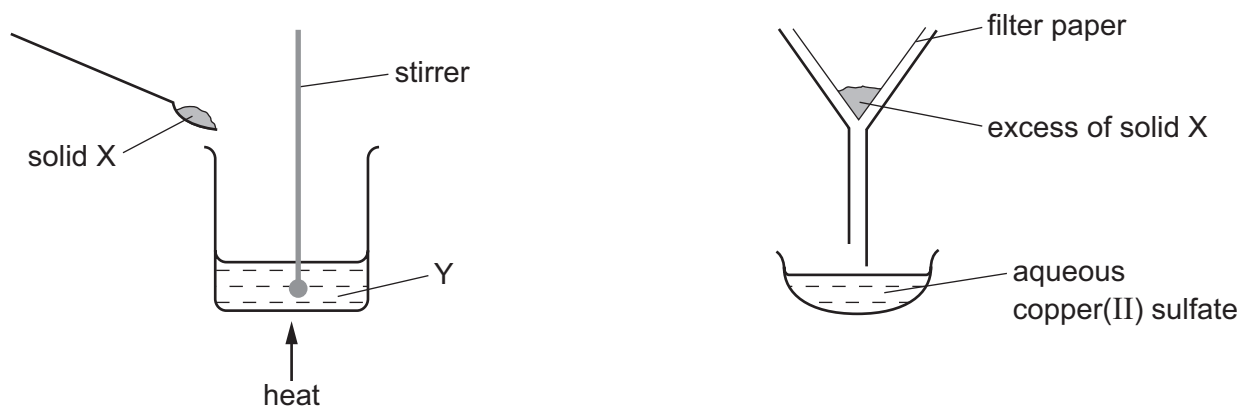
- A acidic
- B amphoteric
- C basic
- D neutral

- 19 Hydrogen chloride gas reacts with water to produce an acidic solution. The equation for the reaction is shown.



Which statement describes what happens during the reaction?

- A** The chloride ion is formed by accepting an electron from the water.  
**B** The hydrogen chloride loses an electron to form the chloride ion.  
**C** The water accepts a proton from the hydrogen chloride.  
**D** The water donates a proton to the hydrogen chloride.
- 20 The apparatus shown is used to prepare aqueous copper(II) sulfate.



What are X and Y?

	X	Y
<b>A</b>	copper	aqueous iron(II) sulfate
<b>B</b>	copper(II) chloride	sulfuric acid
<b>C</b>	copper(II) oxide	sulfuric acid
<b>D</b>	sulfur	aqueous copper(II) chloride

21 Information about some silver compounds is shown in the table.

compound	formula	solubility in water
silver carbonate	$\text{Ag}_2\text{CO}_3$	insoluble
silver chloride	$\text{AgCl}$	insoluble
silver nitrate	$\text{AgNO}_3$	soluble
silver oxide	$\text{Ag}_2\text{O}$	insoluble

Which equation shows a reaction which **cannot** be used to make a silver salt?

- A**  $\text{AgNO}_3(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{HNO}_3(\text{aq})$
- B**  $\text{Ag}_2\text{O}(\text{s}) + 2\text{HNO}_3(\text{aq}) \rightarrow 2\text{AgNO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$
- C**  $\text{Ag}_2\text{CO}_3(\text{s}) + 2\text{HNO}_3(\text{aq}) \rightarrow 2\text{AgNO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$
- D**  $2\text{Ag}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow 2\text{AgCl}(\text{s}) + \text{H}_2(\text{g})$

22 What is **not** a property of Group I metals?

- A** They are soft and can be cut with a knife.
- B** They react when exposed to oxygen in the air.
- C** They produce an acidic solution when they react with water.
- D** They react rapidly with water producing hydrogen gas.

23 Four substances, P, Q, R and S, are tested as shown.

test	substance			
	P	Q	R	S
dilute hydrochloric acid added	gas given off which 'pops' with a lighted splint	gas given off which turns limewater milky	no reaction	no reaction
dilute aqueous sodium hydroxide added and warmed gently	no reaction	no reaction	gas given off which turns damp, red litmus paper blue	no reaction

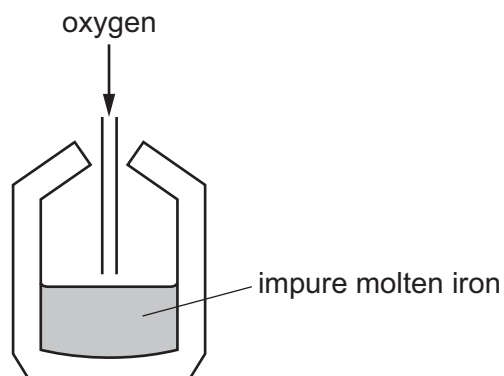
What are P, Q, R and S?

	P	Q	R	S
<b>A</b>	Mg	$\text{Na}_2\text{CO}_3$	$\text{NH}_4\text{Cl}$	$\text{NaCl}$
<b>B</b>	Mg	$\text{NH}_4\text{Cl}$	$\text{Na}_2\text{CO}_3$	$\text{NaCl}$
<b>C</b>	Mg	$\text{Na}_2\text{CO}_3$	$\text{NaCl}$	$\text{NH}_4\text{Cl}$
<b>D</b>	$\text{Na}_2\text{CO}_3$	Mg	$\text{NaCl}$	$\text{NH}_4\text{Cl}$

24 Which statement about transition elements and their compounds is correct?

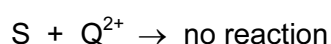
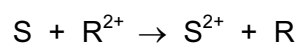
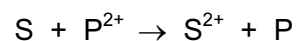
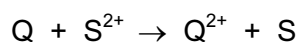
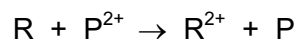
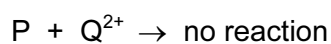
- A** All the transition elements have an oxidation state of +2 only.
- B** Aqueous solutions of the salts of transition elements are generally coloured.
- C** Transition elements change from metal to non-metal across the period.
- D** Transition elements can act as catalysts but their compounds cannot.

25 Impure iron from the blast furnace is converted to steel as shown.



Which statement about the process is correct?

- A Acidic oxides are added to remove alkaline impurities.
  - B Coke is added as a reducing agent.
  - C Oxygen is blown in to oxidise the impure iron.
  - D The steel produced contains less carbon than the impure iron.
- 26 The ionic equations represent the reactions between four metals, P, Q, R and S, and solutions of the salts of the same metals.



What is the correct order of reactivity of the metals?

	most	→			least
<b>A</b>	P	R	S	Q	
<b>B</b>	Q	R	S	P	
<b>C</b>	Q	S	R	P	
<b>D</b>	S	Q	P	R	

27 Aluminium is extracted by electrolysis.

From which ore is aluminium extracted and at which electrode is aluminium deposited during electrolysis?

	ore	electrode
<b>A</b>	bauxite	negative
<b>B</b>	bauxite	positive
<b>C</b>	cryolite	negative
<b>D</b>	cryolite	positive

28 Zinc oxide can be reacted with carbon to produce zinc metal.

Which equation for this reaction is correct?

- A**  $2\text{ZnO} + \text{C} \rightarrow 2\text{Zn} + \text{CO}$
- B**  $2\text{ZnO} + 2\text{C} \rightarrow 2\text{Zn} + 2\text{CO}_2$
- C**  $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$
- D**  $\text{ZnO} + 2\text{C} \rightarrow \text{Zn} + 2\text{CO}_2$

29 Air is a mixture of gases.

Which gas is present in the largest amount?

- A** argon
- B** carbon dioxide
- C** nitrogen
- D** oxygen

30 Which information about carbon dioxide and methane is correct?

		carbon dioxide	methane
<b>A</b>	formed when vegetation decomposes	✓	✗
<b>B</b>	greenhouse gas	✓	✓
<b>C</b>	present in unpolluted air	✗	✗
<b>D</b>	produced during respiration	✗	✓

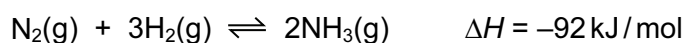
key  
 ✓ = true  
 ✗ = false

31 Underwater steel pipes can be protected from corrosion by attaching magnesium blocks to them.

Which equation represents the reaction that prevents corrosion?

- A  $\text{Fe} \rightarrow \text{Fe}^{2+} + 2\text{e}^{-}$
- B  $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^{-}$
- C  $\text{Mg} \rightarrow \text{Mg}^{2+} + 2\text{e}^{-}$
- D  $\text{Mg}^{2+} + 2\text{e}^{-} \rightarrow \text{Mg}$

32 Ammonia is manufactured by the Haber process. The reaction is exothermic.



Which statement about the Haber process is correct?

- A The reaction is irreversible and produces only one product.
- B The reaction is reversible and produces less ammonia at high pressure.
- C The reaction is reversible and produces less ammonia at high temperature.
- D The reaction is slow because a catalyst is not used in the Haber process.

33 Sulfuric acid is manufactured by the Contact process.

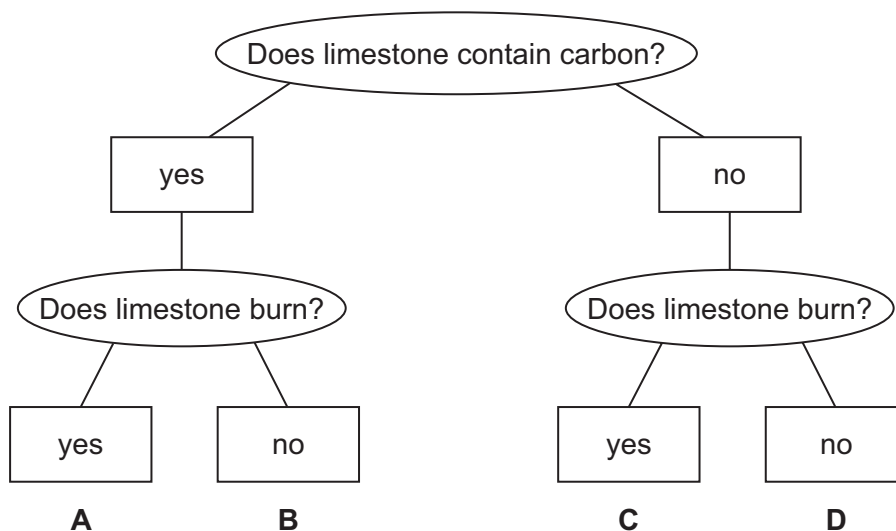
The most important reaction takes place in the presence of a catalyst.

What are the reactants and the catalyst for this reaction?

	reactants	catalyst
<b>A</b>	sulfur and oxygen	vanadium(V) oxide
<b>B</b>	sulfur dioxide and oxygen	vanadium(V) oxide
<b>C</b>	sulfur dioxide and steam	iron
<b>D</b>	sulfur trioxide and water	platinum



34 Which box corresponds to limestone?

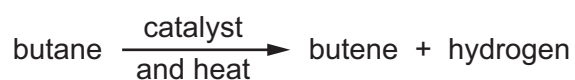


35 Petroleum is an important fossil fuel.

Which row correctly describes petroleum?

	type of substance	composition
<b>A</b>	compound	mainly hydrocarbons
<b>B</b>	compound	only hydrogen and carbon
<b>C</b>	mixture	mainly hydrocarbons
<b>D</b>	mixture	only hydrogen and carbon

36 Butane reacts as shown.



What is this type of reaction?

- A** combustion
- B** cracking
- C** polymerisation
- D** reduction

37 Substance Z has the following characteristics.

- 1 It burns in an excess of oxygen to form carbon dioxide and water.
- 2 It is oxidised by air to form a liquid smelling of vinegar.
- 3 It reacts with carboxylic acids to form esters.

What is substance Z?

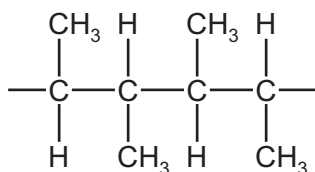
- A ethane
- B ethanoic acid
- C ethanol
- D ethyl ethanoate

38 Ethanol is manufactured by the catalytic addition of steam to ethene and by fermentation.

Which row shows an advantage and a disadvantage of using the catalytic addition of steam to ethene compared to fermentation?

	advantage	disadvantage
<b>A</b>	fast	the product is impure
<b>B</b>	fast	uses non-renewable materials
<b>C</b>	the product is pure	slow
<b>D</b>	uses renewable materials	slow

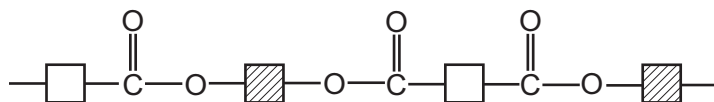
39 The partial structure of addition polymer X is shown.



Which monomer is used to form polymer X?

- A  $\text{CH}_2=\text{CH}_2$
- B  $\text{CH}_3\text{CH}=\text{CH}_2$
- C  $\text{CH}_3\text{CH}=\text{CHCH}_3$
- D  $\text{CH}_3\text{CH}_2\text{CH}=\text{CH}_2$

40 The diagram shows the partial structure of *Terylene*.



From which pair of compounds is it made?

- A**  $\text{HO}-\overset{\text{O}}{\parallel}{\text{C}}-\square-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$  +  $\text{HO}-\square-\text{OH}$
- B**  $\text{HO}-\square-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$  +  $\text{HO}-\square-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$
- C**  $\text{HO}-\square-\text{OH}$  +  $\text{HO}-\overset{\text{O}}{\parallel}{\text{C}}-\square-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$
- D**  $\text{HO}-\overset{\text{O}}{\parallel}{\text{C}}-\square-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$  +  $\text{HO}-\overset{\text{O}}{\parallel}{\text{C}}-\square-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge International Examinations Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at [www.cie.org.uk](http://www.cie.org.uk) after the live examination series.

Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.

### The Periodic Table of Elements

Group																	
I	II											III	IV	V	VI	VII	VIII
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	<b>Key</b> atomic number atomic symbol name relative atomic mass										5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24											1 <b>H</b> hydrogen 1	13 <b>Al</b> aluminium 27	14 <b>Si</b> silicon 28	15 <b>P</b> phosphorus 31	16 <b>S</b> sulfur 32	17 <b>Cl</b> chlorine 35.5
19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	21 <b>Sc</b> scandium 45	22 <b>Ti</b> titanium 48	23 <b>V</b> vanadium 51	24 <b>Cr</b> chromium 52	25 <b>Mn</b> manganese 55	26 <b>Fe</b> iron 56	27 <b>Co</b> cobalt 59	28 <b>Ni</b> nickel 59	29 <b>Cu</b> copper 64	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	39 <b>Y</b> yttrium 89	40 <b>Zr</b> zirconium 91	41 <b>Nb</b> niobium 93	42 <b>Mo</b> molybdenum 96	43 <b>Tc</b> technetium —	44 <b>Ru</b> ruthenium 101	45 <b>Rh</b> rhodium 103	46 <b>Pd</b> palladium 106	47 <b>Ag</b> silver 108	48 <b>Cd</b> cadmium 112	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	54 <b>Xe</b> xenon 131
55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	57–71 lanthanoids	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89–103 actinoids	104 <b>Rf</b> rutherfordium —	105 <b>Db</b> dubnium —	106 <b>Sg</b> seaborgium —	107 <b>Bh</b> bohrium —	108 <b>Hs</b> hassium —	109 <b>Mt</b> meitnerium —	110 <b>Ds</b> darmstadtium —	111 <b>Rg</b> roentgenium —	112 <b>Cn</b> copernicium —	114 <b>Fl</b> flerovium —	116 <b>Lv</b> livermorium —	—	—	—	—

57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175
89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

lanthanoids

actinoids

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.)