

## **MARK SCHEME for the May/June 2015 series**

### **5070 CHEMISTRY**

**5070/32**

Paper 3 (Practical Test), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2015 series for most Cambridge IGCSE<sup>®</sup>, Cambridge International A and AS Level components and some Cambridge O Level components.

<b>Page 2</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>Cambridge O Level – May/June 2015</b>	<b>5070</b>	<b>32</b>

# 1 (a) Titration

Accuracy 8 marks

For the two best titres give:

4 marks for a value within  $0.2\text{ cm}^3$  of supervisor

2 marks for a value within  $0.3\text{ cm}^3$  of supervisor

1 mark for a value within  $0.4\text{ cm}^3$  of supervisor

Concordance 3 marks

Give:

3 marks if all the ticked values are within  $0.2\text{ cm}^3$

2 marks if all the ticked values are within  $0.3\text{ cm}^3$

1 mark if all the ticked values are within  $0.4\text{ cm}^3$

Average 1 mark

Give 1 mark if the candidate calculates a correct average (error not greater than 0.05) of all his/her ticked values.

[12]

## Calculations

Assuming a  $25.0\text{ cm}^3$  pipette and a titre of  $20.2\text{ cm}^3$ .

(b) moles of sodium hydroxide in  $25.0\text{ cm}^3$  of **P**

$$= \frac{25.0 \times 0.0984}{1000}$$

$$= 0.00246$$

[1]

(c) concentration, in  $\text{mol/dm}^3$ , of  $\text{H}_3\text{PO}_3$  in **Q**

$$= \frac{5.04}{82}$$

$$= 0.0615$$

[1]

(d) moles of  $\text{H}_3\text{PO}_3$  in average titre of **Q**

$$= \frac{20.2 \times 0.0615}{1000}$$

$$= 0.00124$$

[1]

<b>Page 3</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>Cambridge O Level – May/June 2015</b>	<b>5070</b>	<b>32</b>

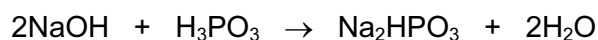
(e) moles of sodium hydroxide which react with 1 mole of  $\text{H}_3\text{PO}_3$

$$= \frac{0.00246}{0.00124}$$

$$= 1.98$$

[1]

(f) balanced equation for the reaction



whole numbers consistent with answer in (e) on left hand side of equation (1)

correct formulae for products and balancing of the equation (1)

[2]

[Total: 18]

**2** R is ammonia S is iron(II) sulfate

Test	Notes
<b>General points</b> For ppt allow solid, suspension, powder  For gases Name of gas requires test to be at least partially correct. Effervesces = bubbles = gas vigorously evolved, but not gas evolved.  Solutions Colourless not equivalent to clear, clear not equivalent to colourless.	
<b>1</b>  gas turns damp red litmus blue (1)  ammonia (1)	to score ammonia mark there must be some indication of a test, i.e. smell of ammonia, alkaline gas, tested with litmus
<b>2</b>  white ppt (1)  soluble in excess (1)  colourless solution (1)	

<b>Page 4</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>Cambridge O Level – May/June 2015</b>	<b>5070</b>	<b>32</b>

<b>Test</b>	<b>Notes</b>
<b>3</b> <b>(a)</b> white ppt (1) <b>(b)</b> solid disappears (1) colourless solution (1)	
<b>4</b> <b>(a)</b> no reaction (1) <b>(b)</b> bubbles (1) gas relights a glowing splint (1) oxygen (1) blue solution (1)	to score oxygen mark there must be some indication of a test, e.g. 'tested with a glowing splint', 'relights a splint'
<b>5</b> <b>(a)</b> white ppt (1) <b>(b)</b> solid remains (1)	
<b>6</b> green ppt (1) insoluble in excess (1) turns brown at surface (1)	

Page 5	Mark Scheme	Syllabus	Paper
	Cambridge O Level – May/June 2015	5070	32

Test	Notes
<b>7</b> <b>(a)</b> filtrate is yellow (1) <b>(b)</b> red-brown ppt (1) insoluble in excess (1)	

Any 20 of the 21 scoring points

[20]

### Conclusions

**R** is ammonia/ $\text{NH}_3$  or ammonium hydroxide/ $\text{NH}_4\text{OH}$   
(ammonia identified in test 1)

(1)

**S** is iron(II) sulfate/ $\text{FeSO}_4$   
(in test 4 white ppt insoluble in acid and in test 6 green ppt)

(1)

[2]

[Total: 22]